**Bohr And Flame Test Quiz – Honors Chemistry**

1. List three ways that the Bohr model differs from the quantum mechanical model of the atom. (6 pt)
2. How did Bohr believe that an atom gave off light when it had energy added to it? Please be complete in your answer. (6 pt)
3. How is spectroscopy used to find the identity of an unknown element? (3 pt)
4. What do the terms “ground state” and “excited state” mean? (4 pt)
5. What’s the difference between line spectra and continuous spectra? (4 pt)
6. Why is the flame test, as we performed it in our lab, not practical in “real world” labs? (3 pt)
7. What are the symbols and names for the four quantum numbers? (4 pt)
8. How does the concept of an electron differ between the Bohr and quantum models of the atom? (4 pt)
9. Write the long and short versions of the electron configuration for phosphorus and sketch its orbital filling diagram. (4 pt)